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[2] (b) Marit arranged the values she found in Table 12 into an energy cycle.  $C_2H_4(g) + H_2(g) \rightarrow C_2H_6(g)$   $3 O_2 + 2 CO_2(g) + 3 H_2O - 1411 kJ mol^{-1} - 286 kJ mol^{-1} - 1560 kJ mol^{-1}$  (hydrogenation) (l) Calculate the value for the enthalpy of hydrogenation of ethene from the energy cycle.....

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8 20 2 2 2 o correct reactants and products; [2] correct coefficients; M2 can only be scored if M1 correct. (b) Local pollutant: carbon monoxide/CO / volatile organics/VOCs / nitrogen oxide/NO / (unburnt) hydrocarbons; Do not accept methane/CH<sub>4</sub>, ethane/C<sub>2</sub>H<sub>6</sub>, propane/C<sub>3</sub>H<sub>8</sub> or butane/C<sub>4</sub>H<sub>10</sub>. Global pollutant: nitrogen oxide/NO / carbon ...

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$CH_3CHO(l) + 3 O_2(g) \rightarrow 2 CO_2(g) + 3 H_2O(l) - 1367 kJ$   $3 C_2H_6 + 2 O_2 \rightarrow 2 C_2H_4 + 2 H_2 = mol^{-1} - 1 CH_3CHO(l) + 2 O_2(g) \rightarrow 2 CO_2(g) + 2 H_2O(l) - 1367 kJ$   $9 HO(l) + 8 O_2 \rightarrow 2 C_2H_6 + 2 O_2 + 2 H_2 = -5470 kJ mol^{-1}$  (i) Calculate the mass, in g, of ethanol and octane in 1.00 kg of the fuel mixture. [1].....

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possible liver damage; [2 max] (b) mild analgesics function by intercepting the pain stimulus at the source / interfere with the production of substances that cause pain/prostaglandins; strong analgesics work by bonding to receptor sites in the brain / prevent the transmission of pain impulses without depressing the central nervous system; [2]

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